

**Honors Chemistry**  
**Extra Stoichiometry Problems**

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.  
a. Write and balance the chemical equation.



- b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many representative particles (and what type) of silver chloride are produced?

$$\frac{39.02 \text{ g BaCl}_2}{1} \times \frac{1 \text{ mole BaCl}_2}{208.233 \text{ g BaCl}_2} \times \frac{2 \text{ moles AgCl}}{1 \text{ mole BaCl}_2} \times \frac{6.02 \times 10^{23} \text{ f. u. AgCl}}{1 \text{ mole AgCl}} = 2.256 \times 10^{23} \text{ f. u. AgCl}$$

- c. If 410.8 grams of barium nitrate are produced how many grams of silver nitrate were reacted?

$$\frac{410.8 \text{ g Ba}(\text{NO}_3)_2}{1} \times \frac{1 \text{ mole Ba}(\text{NO}_3)_2}{261.335 \text{ g Ba}(\text{NO}_3)_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ mole Ba}(\text{NO}_3)_2} \times \frac{169.872 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 534.1 \text{ g AgNO}_3$$

2. Nitrogen gas is reacted with hydrogen gas to form nitrogen trihydride.

- a. Write and balance the chemical equation.



- b. How many liters of nitrogen trihydride are produced at STP if 80.28 grams of hydrogen gas are reacted in an excess of nitrogen?

$$\frac{80.28 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.016 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{22.4 \text{ L NH}_3}{1 \text{ mole NH}_3} = 594.7 \text{ L NH}_3$$

- c. How many grams of hydrogen are needed to fully react 621.9 grams of nitrogen gas?

$$\frac{621.9 \text{ g N}_2}{1} \times \frac{1 \text{ mole N}_2}{28.014 \text{ g N}_2} \times \frac{3 \text{ moles H}_2}{1 \text{ mole N}_2} \times \frac{2.016 \text{ g H}_2}{1 \text{ mole H}_2} = 134.3 \text{ g H}_2$$

- d. How many liters of nitrogen gas are needed to completely react 90.38 L of hydrogen gas at STP?

$$\frac{90.28 \text{ L H}_2}{1} \times \frac{1 \text{ mole H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mole N}_2}{3 \text{ mole H}_2} \times \frac{22.4 \text{ L N}_2}{1 \text{ mole N}_2} = 30.09 \text{ L N}_2$$

$$\frac{90.28 \text{ L H}_2}{1} \times \frac{\text{OR}}{3 \text{ L H}_2} = 30.09 \text{ L N}_2$$

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