Name _____ Date _____ Class _____

CHAPTER 11 STUDY GUIDE

Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write *true* or *false*.

_____ **1.** The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

_____ **2.** Stoichiometry is based on the law of conservation of mass.

_____ **3.** In any chemical reaction, the mass of the products is less than the mass of the reactants.

_____ **4.** The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.

_____ **5.** The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.

methanol + oxygen → carbon dioxide + water

$$2CH_3OH(l) + 3O_2(g) \rightarrow 2CO_2(g) + 4H_2O(g)$$

Substance	Molar Mass (g/mol)	Number of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05			
7. Oxygen gas	32.00			
8. Carbon dioxide	44.01			
9. Water	18.02			

10. What are the reactants? _____

11. What are the products? _____

12. What is the total mass of the reactants? _____

13. What is the total mass of the products? _____

14. How do the total masses of the reactants and products compare? _____

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