

9.6

The reaction is



The variation of $\Delta h_{\text{rxn}}^\circ$ is described by the following equation

$$\Delta h_{\text{rxn}}^\circ = \Delta h_{\text{rxn},298}^\circ + R \left[\Delta A(T - 298) + \frac{\Delta B}{2}(T^2 - 298^2) + \frac{\Delta C}{3}(T^3 - 298^3) - \Delta D \left(\frac{1}{T} - \frac{1}{298} \right) \right]$$

The A, B, C, and D parameters are calculated with data from Table A.2.2:

Species	ν	A	$B \times 10^3$	$C \times 10^6$	$D \times 10^{-5}$
SO ₂	-1	5.699	0.801	0	-1.015
O ₂	-1/2	3.639	0.506	0	-0.227
SO ₃	1	8.06	1.056	0	-2.028

$$\Delta A = -1(5.699) - \frac{1}{2}(3.639) + 1(8.06) = 0.5415$$

$$\Delta B = -1(0.801 \times 10^{-3}) - \frac{1}{2}(0.506 \times 10^{-3}) + 1(1.056 \times 10^{-3}) = 2 \times 10^{-6}$$

$$\Delta C = 0$$

$$\Delta D = -1(-1.015 \times 10^5) - \frac{1}{2}(-0.227 \times 10^5) + 1(-2.028 \times 10^5) = -8.995 \times 10^4$$

From Problem 9.5:

$$\Delta h_{\text{rxn},298}^\circ = -98.96 \times 10^3 \left[\frac{\text{J}}{\text{mol}} \right]$$

Therefore, at 700 °C:

$$\Delta h_{\text{rxn}}^\circ = -98.96 \times 10^3 \left[\frac{\text{J}}{\text{mol}} \right] + \left(8.314 \left[\frac{\text{J}}{\text{mol}} \right] \right) \left[0.5415(T - 298) + \frac{2 \times 10^{-6}}{2}(T^2 - 298^2) - (-8.995 \times 10^4) \left(\frac{1}{T} - \frac{1}{298} \right) \right]$$

Substituting the above expression into Equation 9.51, we get

$$\frac{d \ln K}{dT} = \frac{-98.96 \times 10^3 \left[\frac{\text{J}}{\text{mol}} \right] + \left(8.314 \left[\frac{\text{J}}{\text{mol}} \right] \right) \left[0.5415(T - 298) + \frac{2 \times 10^{-6}}{2}(T^2 - 298^2) - (-8.995 \times 10^4) \left(\frac{1}{T} - \frac{1}{298} \right) \right]}{RT^2}$$

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