

ACID-BASE TITRATION

Name _____

An Acid-Base Titration is the determination of the concentration of an acid or base by exactly neutralizing the acid or base with an acid or base of known concentration.

Example: It takes 45.2 ml of 0.1 M hydrochloric acid to neutralize 30 ml of sodium hydroxide. What was the concentration of sodium hydroxide?



- Step 1: Express the known concentration in liters: .045 L of HCl
Step 2: Multiply by its concentration: (.045 L)(0.1 M) = .0045
Step 3: Multiply by the mole to mole comparison (see equation)
(.0045) (1/1) = .0045 still
Step 4: Divide by the unknown concentration in liters
.0045 / .030 L = .150 M

1. In the titration of 10 ml of vinegar 40 ml of .2 M sodium hydroxide was used. What was the molarity of the vinegar? $\text{NaOH} + \text{CH}_3\text{COOH} \rightarrow \text{H}_2\text{O} + \text{CH}_3\text{COONa}$

2. If 40 ml of .1 M sodium hydroxide neutralizes 30 ml of HCl, what was the molarity of the HCl? $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

3. It took 1.8 ml of .2 molarity to titrate 45 ml of $\text{Ca}(\text{OH})_2$. What was the molarity of the $\text{Ca}(\text{OH})_2$? $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$

4. 50 ml of .5 M of HCl is needed to neutralize 50 ml of NaOH. Calculate the molarity of the NaOH. $\text{NaOH} + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{NaCl}$

5. To titrate 30 ml of potassium hydroxide required 20 ml of .1 M Hydrochloric Acid. Write the equation and tell the molarity of the potassium hydroxide.

Acid Base Titration Virtual Lab Worksheet Answers

DP Hallahan



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