

Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

Property	Acids	Bases
Conductivity in Aqueous Solution	Very	Very
Taste	Sour	Bitter
Feel	—	Slippery
Reaction with Metal	produces H_2 gas	—
Reaction with Carbonate	produces CO_2 gas	—
Colour on Litmus Paper	Red	Blue
Colour with Universal Indicator	Red	Purple
Colour with phenolphthalein	Colourless	pink

2. Indicate if each compound is an acid or base and name it.

	Acid/Base?	Name
HNO_3	<u>A</u>	<u>nitric acid</u>
$Mg(OH)_2$	<u>B</u>	<u>magnesium hydroxide</u>
H_3PO_4	<u>A</u>	<u>phosphoric acid</u>
H_2CO_3	<u>A</u>	<u>carbonic acid</u>
HF	<u>A</u>	<u>hydrofluoric acid</u>
$Fe(OH)_3$	<u>B</u>	<u>iron (III) hydroxide</u>

3. Write the formula for each of the following acids/bases:

hydrochloric acid	<u>HCl</u>	hydrochloric acid	<u>HCl</u>
aluminum hydroxide	<u>$Al(OH)_3$</u>	sodium hydroxide	<u>NaOH</u>
lithium hydroxide	<u>LiOH</u>	calcium hydroxide	<u>$Ca(OH)_2$</u>
sulfuric acid	<u>H_2SO_4</u>	sulfurous acid	<u>H_2SO_3</u>

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Table of Contents Acid Base Worksheet Answer Key

1. Understanding the eBook Acid Base Worksheet Answer Key
 - The Rise of Digital Reading Acid Base Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Worksheet Answer Key
 - Personalized Recommendations

- Acid Base Worksheet Answer Key User Reviews and Ratings
 - Acid Base Worksheet Answer Key and Bestseller Lists
5. Accessing Acid Base Worksheet Answer Key Free and Paid eBooks
 - Acid Base Worksheet Answer Key Public Domain eBooks
 - Acid Base Worksheet Answer Key eBook Subscription Services
 - Acid Base Worksheet Answer Key Budget-Friendly Options
 6. Navigating Acid Base Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid Base Worksheet Answer Key Compatibility with Devices
 - Acid Base Worksheet Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid Base Worksheet Answer Key
 - Highlighting and Note-Taking Acid Base Worksheet Answer Key
 - Interactive Elements Acid Base Worksheet Answer Key
 8. Staying Engaged with Acid Base Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Worksheet Answer Key
 9. Balancing eBooks and Physical Books Acid Base Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Worksheet Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Acid Base Worksheet Answer Key
 - Setting Reading Goals Acid Base Worksheet Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Acid Base Worksheet Answer Key
 - Fact-Checking eBook Content of Acid Base Worksheet Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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