

Name \_\_\_\_\_

## 5-3 Review and Reinforcement

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### Periodic Trends

Use the periodic table and your knowledge of periodic trends to answer the following questions.

Which atom in each pair has the larger atomic radius?

- |           |             |
|-----------|-------------|
| <u>K</u>  | 1. Li or K  |
| <u>Ca</u> | 2. Ca or Na |
| <u>Ca</u> | 3. Ga or B  |
| <u>O</u>  | 4. O or C   |
| <u>Br</u> | 5. Cl or Br |
| <u>Ba</u> | 6. Be or Ba |
| <u>S</u>  | 7. Si or S  |
| <u>Au</u> | 8. Fe or Au |

Which ion in each pair has the smaller atomic radius?

- |                             |                            |
|-----------------------------|----------------------------|
| <u><math>K^+</math></u>     | 9. $K^+$ or $Ca^{2+}$      |
| <u><math>Ba^{2+}</math></u> | 10. $Ba^{2+}$ or $F^-$     |
| <u><math>Al^{3+}</math></u> | 11. $Al^{3+}$ or $P^{3-}$  |
| <u><math>K^+</math></u>     | 12. $K^+$ or $Ca^{2+}$     |
| <u><math>Fe^{3+}</math></u> | 13. $Fe^{2+}$ or $Fe^{3+}$ |
| <u><math>F^-</math></u>     | 14. $F^-$ or $O^{2-}$      |

Which atom or ion in each pair has the larger ionization energy?

- |           |              |
|-----------|--------------|
| <u>O</u>  | 15. Na or O  |
| <u>Ba</u> | 16. Ba or Be |
| <u>F</u>  | 17. Ar or F  |
| <u>Cu</u> | 18. Cu or Fe |
| <u>Ni</u> | 19. I or Ni  |
| <u>V</u>  | 20. K or V   |
| <u>Ca</u> | 21. Ca or Fr |
| <u>Se</u> | 22. W or Se  |

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