

Organic Chemistry

1. Which is not a structural isomer of $\text{CH}_3\text{CH}_2\text{COOH}$?

- A. $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B. $\text{CH}_3\text{COOCH}_3$
- C. $\text{CH}_3\text{COCH}_2\text{OH}$
- D. $\text{HOCH}=\text{CHOCH}_3$
- E. $\text{CH}_3\text{CH}(\text{OH})\text{CHO}$

2. Which is a structural isomer of $\text{CH}_3\text{CH}_2\text{COOH}$ that has a lower boiling point?

- A. $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B. $\text{CH}_3\text{COOCH}_3$
- C. $\text{CH}_3\text{COCH}_2\text{OH}$
- D. $\text{HOCH}=\text{CHOCH}_3$
- E. $\text{CH}_3\text{CH}(\text{OH})\text{CHO}$

3. Which is a structural isomer of $\text{CH}_3\text{CH}_2\text{COOH}$ that exhibits optical isomerism?

- A. $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B. $\text{CH}_3\text{COOCH}_3$
- C. $\text{CH}_3\text{COCH}_2\text{OH}$
- D. $\text{HOCH}=\text{CHOCH}_3$
- E. $\text{CH}_3\text{CH}(\text{OH})\text{CHO}$

4. Which is a structural isomer of $\text{CH}_3\text{CH}_2\text{COOH}$ that exhibits geometric isomerism?

- A. $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B. $\text{CH}_3\text{COOCH}_3$
- C. $\text{CH}_3\text{COCH}_2\text{OH}$
- D. $\text{HOCH}=\text{CHOCH}_3$
- E. $\text{CH}_3\text{CH}(\text{OH})\text{CHO}$

5. Which is a structural isomer of $\text{CH}_3\text{CH}_2\text{COOH}$ that does not exhibit optical isomerism and has the following IR spectrum:



- A. $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B. $\text{CH}_3\text{COOCH}_3$
- C. $\text{CH}_3\text{COCH}_2\text{OH}$
- D. $\text{HOCH}=\text{CHOCH}_3$
- E. $\text{CH}_3\text{CH}(\text{OH})\text{CHO}$

6. Which of the following has four stereoisomers?



7. Which of the following compounds is a structural isomer of one or more of the other compounds and also exhibits geometrical isomerism?

- A. $\text{H}_2\text{C}=\text{CHCH}_2\text{CH}_2\text{CH}_3$
- B. $\text{CH}_3\text{CH}=\text{CHCH}_2\text{CH}_3$
- C. $\text{H}_2\text{C}=\text{CHCH}=\text{CHCH}_3$
- D. $(\text{CH}_3)_2\text{C}=\text{CHCH}_3$

8. Which of the following has the lowest vapor pressure at 25 °C.



9. Which liquid would you expect to have the highest viscosity at room temperature?

- A. n-octane
- B. isooctane
- C. 2-octanone
- D. octanoic acid

10. Predict the length of each NO bond in nitrobenzene ($\text{C}_6\text{H}_5\text{NO}_2$). [The NO single bond length is 146 pm, the N=O double bond length is 121 pm.]

- A. 110 pm
- B. 121 pm
- C. 130 pm
- D. 146 pm

11. According to the IUPAC rules, the following compound can be named:



- A. 1,3-cycloheptadiene
- B. 3,5-cyclohexadiene

20ap Chemistry Multiple Choice Exam

Ying Liu



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