

AP Multiple Choice Questions.

1989

27.

- I. Difference in temperature between freezing point of solvent and freezing point of solution
- II. Molal freezing point depression constant,  $K_f$ , for solvent

In addition to the information above, which of the following gives the minimum data required to determine the molecular mass of a nonionic substance by the freezing point depression technique?

- a. No further information is necessary.
- b. Mass of solute
- c. Mass of solute and mass of solvent
- d. Mass of solute and volume of solvent
- e. Mass of solute, mass of solvent, and vapor pressure of solvent

71. A solution of toluene (molecular weight 92.1) in benzene (molecular weight 78.1) is prepared. The mole fraction of toluene in the solution is 0.100. What is the molality of the solution?

- a. 0.100 m
- b. 0.703 m
- c. 0.921 m
- d. 1.28 m
- e. 1.42 m

1994

14. Which of the following is lower for a 1.0-molar aqueous solution of any solute than it is for pure water?

- a. pH
- b. Vapor pressure
- c. Freezing point
- d. Electrical conductivity
- e. Absorption of visible light

44. Which of the following solutions has the lowest freezing point?

2002

26. Approximately what mass of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  (250 g mol<sup>-1</sup>) is required to prepare 250 mL of 0.10 M copper(II) sulfate solution?

(A) 4.0 g (B) 6.2 g (C) 34 g (D) 85 g (E) 140 g

38. A 0.10 M aqueous solution of sodium sulfate,  $\text{Na}_2\text{SO}_4$ , is a better conductor of electricity than a 0.10 M aqueous solution of sodium chloride,  $\text{NaCl}$ . Which of the following best explains this observation?

- a.  $\text{Na}_2\text{SO}_4$  is more soluble in water than  $\text{NaCl}$  is.
- b.  $\text{Na}_2\text{SO}_4$  has a higher molar mass than  $\text{NaCl}$  has.
- c. To prepare a given volume of 0.10 M solution, the mass of  $\text{Na}_2\text{SO}_4$  needed is more than twice the mass of  $\text{NaCl}$  needed.
- d. More moles of ions are present in a given volume of 0.10 M  $\text{Na}_2\text{SO}_4$  than in the same volume of 0.10 M  $\text{NaCl}$ .
- e. The degree of dissociation of  $\text{Na}_2\text{SO}_4$  in solution is significantly greater than that of  $\text{NaCl}$ .

45. What is the mole fraction of ethanol,  $\text{C}_2\text{H}_5\text{OH}$ , in an aqueous solution that is 46 percent ethanol by mass? (The molar mass of  $\text{C}_2\text{H}_5\text{OH}$  is 46 g; the molar mass of  $\text{H}_2\text{O}$  is 18 g.)

- a. 0.25

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**RJ Shavelson**



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## **Table of Contents Ap Chemistry Multiple Choice Questions**

1. Understanding the eBook Ap Chemistry Multiple Choice Questions
  - The Rise of Digital Reading Ap Chemistry Multiple Choice Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Multiple Choice Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chemistry Multiple Choice Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Multiple Choice Questions
  - Personalized Recommendations
  - Ap Chemistry Multiple Choice Questions User Reviews and Ratings
  - Ap Chemistry Multiple Choice Questions and Bestseller Lists
5. Accessing Ap Chemistry Multiple Choice Questions Free and Paid eBooks
  - Ap Chemistry Multiple Choice Questions Public Domain eBooks
  - Ap Chemistry Multiple Choice Questions eBook Subscription Services
  - Ap Chemistry Multiple Choice Questions Budget-Friendly Options

6. Navigating Ap Chemistry Multiple Choice Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Ap Chemistry Multiple Choice Questions Compatibility with Devices
  - Ap Chemistry Multiple Choice Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ap Chemistry Multiple Choice Questions
  - Highlighting and Note-Taking Ap Chemistry Multiple Choice Questions
  - Interactive Elements Ap Chemistry Multiple Choice Questions
8. Staying Engaged with Ap Chemistry Multiple Choice Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chemistry Multiple Choice Questions
9. Balancing eBooks and Physical Books Ap Chemistry Multiple Choice Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chemistry Multiple Choice Questions
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Multiple Choice Questions
  - Setting Reading Goals Ap Chemistry Multiple Choice Questions
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Multiple Choice Questions
  - Fact-Checking eBook Content of Ap Chemistry Multiple Choice Questions
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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