

Chemical Reactions: Basic Stoichiometry PhET Lab

<https://phet.colorado.edu/en/simulation/reactants-products-and-leftovers>

The conservation of atoms in a chemical reaction leads to the ability to calculate the amount of products formed and reactants used (stoichiometry).

Introduction:

When we bake/cook something, we use a specific amount of each ingredient. Imagine if you made a batch of cookies and used way too many eggs, or not enough sugar. YUCK! In chemistry, reactions proceed with very specific recipes. The study of these recipes is *stoichiometry*. When the reactants are present in the correct amounts, the reaction will produce products. What happens if there are more or less of some of the reactants present?

Vocabulary: Before you begin using the simulation, please define the following:

Stoichiometry: _____

Limiting Reactant: _____

Excess Reactant: _____

Synthesis Reaction: _____

Combustion Reaction: _____

Mole Ratio: _____

Diatomic Molecule: _____

Mole: _____

Hydrocarbon: _____

Procedure: PhET Simulations → Play with the Sims → Chemistry → Reactants, Products, and Leftovers View Next!

If a yellow bar drops down in your browser, click on it and select "Allow Blocked Content"

PART 1: MAKING SANDWICHES: Sandwich Shop

1. The 🍞 Cheese Sandwich is a simulation of a two-reactant synthesis reaction. In this case, one reactant will be *limiting*, while the other will be in *excess*.
2. Take some time and familiarize yourself with the simulation.
3. Set the reaction to a simple mole ratio of 2:1:1



Basic Stoichiometry Phet Lab Worksheet

LL Leslie



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