

Topic 1 Molecules, transport and health

1A Chemistry for biologists

1A Checkpoint

1A.1 The chemistry of life

- In some covalent compounds, the electrons in the covalent bonds are not quite evenly shared. This means the molecule has a part that is slightly negative and a part that is slightly positive. This separation of charge is called a dipole, and the tiny charges are represented as δ^+ and δ^- . A molecule containing dipoles is known as a polar molecule. Dipoles are particularly common if one or more hydrogen atoms are involved in the bond.
- Ionic substances: formed when atoms are joined by ionic bonds. The positive or negative ions formed when the atoms lose or gain electrons form an ionic compound.

Polar substances: compounds formed with covalent bonds where there is a slightly unequal sharing of the electrons across the bond, giving a covalent molecule with slight dipoles that can affect intermolecular bonding.
- Water is a polar molecule because the electrons are held closer to the oxygen atom than to the hydrogen atoms, so the oxygen has a very small negative charge and the hydrogen atoms have very small positive charges. As a result, the slightly negative region of one water molecule is attracted to the slightly positive region of another water molecule, and this weak electrostatic attraction is a hydrogen bond. Therefore, water has relatively high melting and boiling points and solid water is less dense than liquid water.
- Excellent solvent: its covalent nature means water dissolves covalent compounds, its polar nature means it dissolves ionic compounds, so it is an ideal solvent for chemical reactions in biological systems and makes an excellent transport medium.

It is slow to absorb and release heat so is a very stable medium for life.

It is a liquid at room temperature so cannot be compressed, therefore it can be used in hydraulic systems.

Water molecules are cohesive (stick together) and adhesive (stick to other substances). Both are important properties in the movement of water up the xylem of plants.

High surface tension means that the water surface acts as a skin; this is very important in water being drawn up the phloem of a plant and for life at the surface of ponds, lakes and other water masses.

Also award marks for any other sensible points.

1A.2 Carbohydrates 1: monosaccharides and disaccharides

- Carbohydrates are important in cells as a usable energy source. They are also important for storing energy, and in plants, fungi and bacteria they form an important part of the cell wall. The basic structure of all carbohydrates is the same. They are made up of carbon, hydrogen and oxygen. There are three main groups of carbohydrates with varying complexity of molecules: monosaccharides, disaccharides and polysaccharides.

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