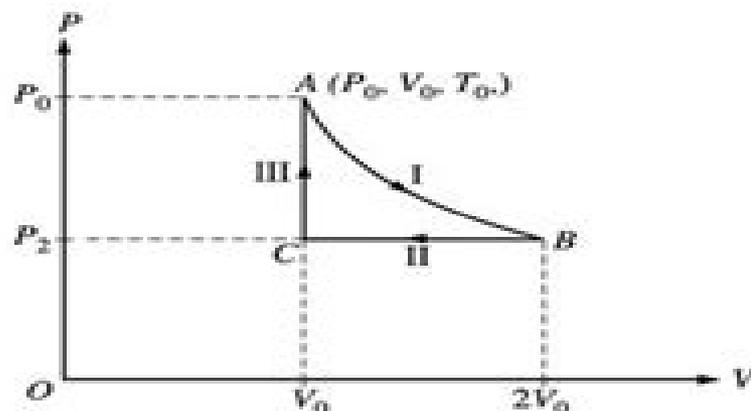


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5. (10 points)

A sample of ideal gas is taken through steps I, II, and III in a closed cycle, as shown on the pressure  $P$  versus volume  $V$  diagram above, so that the gas returns to its original state. The steps in the cycle are as follows.

I. An isothermal expansion occurs from point  $A$  to point  $B$ , and the volume of the gas doubles.

II. An isobaric compression occurs from point  $B$  to point  $C$ , and the gas returns to its original volume.

III. A constant volume addition of heat occurs from point  $C$  to point  $A$  and the gas returns to its original pressure.

(a) Determine numerical values for the following ratios, justifying your answers in the spaces next to each ratio.

i.  $\frac{P_B}{P_A} =$

ii.  $\frac{P_C}{P_A} =$

iii.  $\frac{T_B}{T_A} =$

iv.  $\frac{T_C}{T_A} =$

(b) During step I, the change in internal energy is zero. Explain why.

(c) During step III, the work done on the gas is zero. Explain why.

# Answers To Ap Physics B Work 1

**Yicheng Fang**



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