

# Covalent Formulas Worksheet

Name: \_\_\_\_\_

## Covalent Compounds- Writing Formulas



**Background:** Covalent compounds are formed when two nonmetals share electrons. Instead of charges, we use prefixes in the compound's name to tell us how many atoms of each element are present.

The first element in the name keeps its regular name. The second element changes its ending to *-ide*.

### Example

- Carbon dioxide → CO<sub>2</sub>
- Prefix "di-" = 2 oxygens. No prefix on carbon → 1 carbon.



**Directions:** Write the correct formula for each covalent compound.

1) Carbon monoxide

4) Nitrogen dioxide

2) Dinitrogen trioxide

3) Carbon tetrachloride

a) CO

d) CCl<sub>4</sub>

b) N<sub>2</sub>O<sub>3</sub>

c) N<sub>2</sub>O<sub>2</sub>



\_\_\_\_\_ To name

\_\_\_\_\_ of an *-ide*.



## Physical Science Chemistry

# Covalent Naming Worksheet 233 Physical Science

**O García**

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