

## Unit 1 Chemistry Study Guide

### History of the Atom:

#### Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

#### Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

#### Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

#### Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

#### John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
  - o All matter is composed of extremely small particles called atoms.
  - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
  - o Atoms cannot be subdivided, created, or destroyed.
  - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
  - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
  - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

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**Arshad Iqbal**



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## **Table of Contents Chemistry Study Guide Packet Answer**

1. Understanding the eBook Chemistry Study Guide Packet Answer
  - The Rise of Digital Reading Chemistry Study Guide Packet Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Packet Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Study Guide Packet Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Packet Answer
  - Personalized Recommendations
  - Chemistry Study Guide Packet Answer User Reviews and Ratings
  - Chemistry Study Guide Packet Answer and Bestseller Lists
5. Accessing Chemistry Study Guide Packet Answer Free and Paid eBooks
  - Chemistry Study Guide Packet Answer Public Domain eBooks
  - Chemistry Study Guide Packet Answer eBook Subscription Services
  - Chemistry Study Guide Packet Answer Budget-Friendly Options

6. Navigating Chemistry Study Guide Packet Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Study Guide Packet Answer Compatibility with Devices
  - Chemistry Study Guide Packet Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Study Guide Packet Answer
  - Highlighting and Note-Taking Chemistry Study Guide Packet Answer
  - Interactive Elements Chemistry Study Guide Packet Answer
8. Staying Engaged with Chemistry Study Guide Packet Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Study Guide Packet Answer
9. Balancing eBooks and Physical Books Chemistry Study Guide Packet Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Study Guide Packet Answer
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide Packet Answer
  - Setting Reading Goals Chemistry Study Guide Packet Answer
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide Packet Answer
  - Fact-Checking eBook Content of Chemistry Study Guide Packet Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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