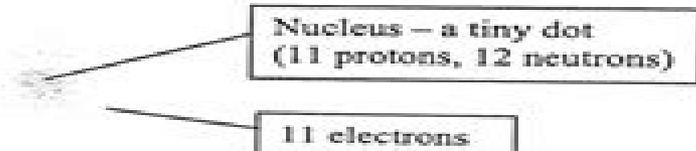
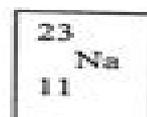


## Subatomic Particles

Particle	Symbol	Relative Charge	Absolute Mass	Relative Mass
electron	$e^-$	-1	$9.109 \times 10^{-31}$ kg	0
proton	$p^+$	+1	$1.673 \times 10^{-27}$ kg	1
neutron	$n^0$	0	$1.675 \times 10^{-27}$ kg	1

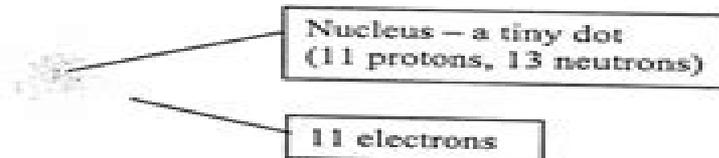
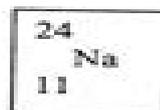
**Model: Two Isotopes of Sodium**

The diagrams below show representations of sodium isotopes. [Note: the diameter of an atom is about 10,000 times larger than the diameter of the atomic nucleus so the relative sizes of the atom and the nucleus are not accurately depicted in these diagrams.]

**Isotope 1**

Nucleus – a tiny dot  
(11 protons, 12 neutrons)

11 electrons

**Isotope 2**

Nucleus – a tiny dot  
(11 protons, 13 neutrons)

11 electrons

**Key Questions**

1. What information is provided by the atomic number, Z?

# of protons

2. What information is provided by the mass number, A?

protons + neutrons

# Chemistry Unit 2 Packet Key

**Rachel Sandford**



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