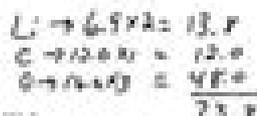


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Mr. A

11. a. How many moles are in 125g of Li_2CO_3 ?

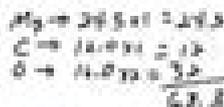
$$125g \text{ Li}_2\text{CO}_3 \times \frac{1 \text{ mol Li}_2\text{CO}_3}{75.8 \text{ g Li}_2\text{CO}_3} = \boxed{1.64 \text{ mol}}$$

- b. If the moles from part a are added to 2.5L of water, what is the molarity?

$$M = \frac{\text{mol}}{L} = \frac{1.64}{2.5} = \boxed{0.656 \text{ M}}$$

12. a. What is the mass of 3.9 moles of MgSO_4 ?

$$3.9 \text{ mol} \times \frac{60.3 \text{ g}}{1 \text{ mol}} = \boxed{235.17 \text{ g}}$$



- b. What would the density (in g/L) of the solution from part a be if it has a volume of 190.2 mL?

$$D = \frac{m}{V} = \frac{235.17}{190.2} = \boxed{1.236 \text{ g/mL}}$$

13. 19.2 moles of hydrochloric acid react, how many grams of hydrogen gas at standard conditions are formed?

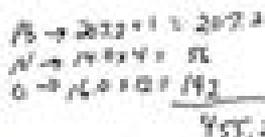
$$19.2 \text{ mol HCl} \times \frac{2 \text{ mol H}_2}{4 \text{ mol HCl}} = 9.6 \text{ mol H}_2 \times \frac{2.016 \text{ g H}_2}{1 \text{ mol H}_2} = \boxed{19.35 \text{ g}}$$

14. The reaction in which sulfur and chlorine react to make sulfur tetrachloride is given below. How many moles of sulfur tetrachloride are produced by the reaction of 8.2 moles of chlorine gas?

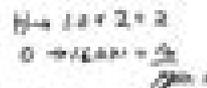
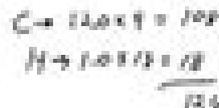
$$8.2 \text{ mol Cl}_2 \times \frac{1 \text{ mol SCl}_4}{4 \text{ mol Cl}_2} = \boxed{2.05 \text{ mol SCl}_4}$$

15. What mass of potassium nitrate will be produced if 100.2 g of lead (Pb) nitrate is reacted with excess potassium phosphate according to the reaction listed below?

$$100.2 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{327.2 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol KNO}_3}{3 \text{ mol Pb(NO}_3)_2} \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = \boxed{20.2 \text{ g KNO}_3}$$



$$\begin{aligned}
 8.2 \text{ mol KNO}_3 \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} &= \boxed{829.02 \text{ g}} \\
 \text{K} &\rightarrow 39.10 \times 3 = 117.30 \\
 \text{N} &\rightarrow 14.01 \times 3 = 42.03 \\
 \text{O} &\rightarrow 16.00 \times 6 = 96.00 \\
 &\hline
 &255.33
 \end{aligned}$$



Name _____

16. Balance these with oxygen according to the following reaction:



- a. If 301.2 g of C_2H_6 was used, how many grams of H_2O could be produced?

$$301.2 \text{ g} \times \frac{1 \text{ mol C}_2\text{H}_6}{74.12 \text{ g C}_2\text{H}_6} = 4.06 \text{ mol C}_2\text{H}_6 \times \frac{6 \text{ mol H}_2\text{O}}{2 \text{ mol C}_2\text{H}_6} = 12.18 \text{ mol H}_2\text{O}$$

$$12.18 \text{ mol H}_2\text{O} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{219.24 \text{ g H}_2\text{O}}$$

- b. If 690.5 L of O_2 under normal conditions was used, how many grams of H_2O could be produced?

$$690.5 \text{ L O}_2 \times \frac{1 \text{ mol O}_2}{22.4 \text{ L O}_2} = 30.8 \text{ mol O}_2 \times \frac{6 \text{ mol H}_2\text{O}}{7 \text{ mol O}_2} = 26.4 \text{ mol H}_2\text{O}$$

$$26.4 \text{ mol H}_2\text{O} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{475.2 \text{ g H}_2\text{O}}$$

Also theoretical for 240 (160 + 80) it's not here

- c. Using the information from part a, what was the limiting reagent?



- d. Using the amount produced by your limiting reagent as your theoretical yield, what would the percentage yield be if your experimental yield was 374.4 g of H_2O ?

$$\frac{\text{Exp. yield}}{\text{Theor. yield}} \times 100 = \frac{374.4}{475.2} \times 100 = \boxed{78.8\%}$$

- e. Using the amount produced by your limiting reagent as your theoretical yield, what would the percentage yield be if your experimental yield was 274.4 g of H_2O ?

$$\frac{\text{Exp. yield}}{\text{Theor. yield}} \times 100 = \frac{274.4}{475.2} \times 100 = \boxed{57.7\%}$$

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Huangqi Zhang



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