

# AP Chemistry Unit 7 Review

## Acids and Bases

Honors Lecture Notes: Unit 7

### Definitions

An Arrhenius acid is a  $H^+$  donor, and an Arrhenius base is an  $OH^-$  donor. These definitions are not always accurate, though, though, and so instead we use the Bronsted-Lowry definitions. A Bronsted-Lowry acid is the same as an Arrhenius acid- a proton donor. However, a Bronsted-Lowry base is a proton acceptor. By accepting protons from water, hydroxide ions are created, causing the solution to be basic.

Substances that can act as either acids or bases (such as water) are called amphoteric. Whether an amphoteric substance acts as an acid or as a base depends on what it is combined with.

A conjugate pair is two species that differ from each other by a single proton. Examples:  $NH_3/NH_4^+$ ,  $HClO_2/ClO_2^-$ . The stronger the acid or base is, the weaker the conjugate is.

### Ions In Solution

Typically, most anions create basic solutions, as due to their negative charge they will accept protons from water readily. The exception to this rule is the conjugates of the strong acids. As strong acids dissociate completely, their conjugate ions tend to be ineffective bases. The strong acids are:  $HCl$ ,  $HNO_3$ ,  $H_2SO_4$ ,  $HI$ ,  $HBr$ , and  $HClO_4$ .

Ex:  $Na_2CO_3$  is basic in solution because:



$NaCl$ , however, is neutral in solution, as  $Cl^-$  is the conjugate of  $HCl$  and thus will not accept protons from water.

Unlike anions, cations are generally not considered to be acidic as they don't often have protons to donate (notable exceptions:  $NH_4^+$ ). Strong bases are defined as those which donate 100% of their  $OH^-$  ions in solution. Strong bases are any alkali hydroxide as well as  $Ba(OH)_2$  and  $Sr(OH)_2$ .

### pH/pOH

To determine the relative acidity/basicity of a solution, the pH/pOH is used.  $pH = -\log[H^+]$  and  $pOH = -\log[OH^-]$ . At  $25^\circ C$ ,  $pH + pOH = 14$  and  $[H^+][OH^-] = 1.0 \times 10^{-14}$  for any solution. Thus, as  $[H^+]$  goes up, pH goes down. Likewise, as pH goes down, pOH goes up.

Any solution with a pH of less than 7 is considered acidic, and a pH of above 7 indicates a basic solution. Note that pH/pOH are NOT limited to a 0-14 scale. Sufficiently concentrated strong acids can have negative pHs, and concentrated strong bases can have pHs in excess of 14.

### Auto-ionization of Water



For water at  $25^\circ C$ ,  $K_w = 1.0 \times 10^{-14}$ , and  $[H^+] = [OH^-] = 1.0 \times 10^{-7}$  M. The  $K_w$  of water (like any equilibrium constant) changes with temperature. The auto-ionization of water is an endothermic process, and as such, if temperature increases  $K_w$  increases. This means that at temperatures above  $25^\circ C$ , the pH of pure water is NOT 7.0. However, neither is the pOH. For instance, at  $50^\circ C$ ,  $K_w = 5.5 \times 10^{-14}$ , and thus  $[H^+] = [OH^-] = 2.3 \times 10^{-7}$  M. At  $50^\circ C$ ,  $pH = pOH = 6.63$ , and  $pH + pOH = 13.2$ .

### $K_a/K_b$

The strength of a weak acid or base is defined by the equilibrium constant. These equilibrium constants are calculated the same way as any other equilibrium constant [products/reactants], but as it is always the loss/gain of a single proton, there are never any exponents.  $K_a/K_b$  values are often used to calculate the pH of a solution containing a weak acid or base.

Ex: Calculate the pH of 1.5 M acetic acid,  $HC_2H_3O_2$ . ( $K_a = 1.8 \times 10^{-5}$ ).



$$K_a = \frac{[H^+][C_2H_3O_2^-]}{[HC_2H_3O_2]}$$

As  $H^+$  and  $C_2H_3O_2^-$  are in a 1:1 ratio, both values can be replaced by  $x$ . So:  $1.8 \times 10^{-5} = \frac{x^2}{1.5}$

$$x = [H^+] = 0.0051 \text{ M} \quad -\log(0.0051 \text{ M}) = 2.28$$

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**Christian G. Meyer**



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web pedigree analysis activity answer key q1 d is the correct answer q2 b is the correct answer why because people affected by this disease are surviving at least to the age they are able to have children q3 c is the correct answer q4

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web pedigree analysis key takeaways pedigrees can help us to analyze the inheritance patterns of many traits pedigrees are typically used in the setting of genetic disorders the most common inheritance patterns include autosomal recessive autosomal dominant and x linked recessive some other less common inheritance patterns include x linked

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web interpreting a human pedigree use the pedigree below to answer 1 in a pedigree a square represents a male if it is darkened he has hemophilia if clear he had normal blood clotting a how many males are there 8 b how many males have hemophilia 3 a circle represents a female if it is darkened she has hemophilia if open she is normal

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web the conclusions about inheritance above can be used to help analyze pedigrees for each pedigree below tell if the trait can be autosomal dominant autosomal recessive and x linked recessive if the pedigree cannot fit a mode of inheritance tell why record your answers in the table on the next page pedigree a pedigree b 2

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web this blog post is going to explain how to solve pedigree problems easily this lecture explains about the different rules of pedigree analysis it explains how to find a pedigree based on characteristics with examples as dominant pedigree recessive pedigree and x linked pedigree x linked recessive trait male offspring are affected more 1

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