

## Stoichiometry Worksheet 2 Answer Key

1.

- a. 2 / 13
- b. 13 / 8
- c. 13 / 10
- d. 2 / 8 (or 1 / 4)
- e. 2 / 10 (or 1 / 5)

2. The  $\text{KClO}_3 / \text{O}_2$  molar ratio is 2/3.

$$2 \text{ mol KClO}_3 / 3 \text{ mol O}_2 = 12.00 \text{ mol KClO}_3 / x = 18.00 \text{ mol.}$$

$$x = 18.00 \text{ mol of O}_2$$

3. Given the following equation:  $2 \text{ K} + \text{Cl}_2 \rightarrow 2 \text{ KCl}$

How many grams of KCl is produced from 2.50 g of K and excess  $\text{Cl}_2$ .

$\frac{2.50 \text{ g of K}}{39.098 \text{ g/mol}}$	=	0.0639 mol of K
$\frac{\text{K}}{2}$	=	$\frac{0.0639 \text{ mol of K}}{x}$
$\frac{\text{KCl}}{2}$	=	$\frac{x}{x}$
		$x = 0.0639 \text{ mol of KCl}$
$0.0639 \text{ mol} \times 74.551 \text{ g/mol of KCl} = 4.77 \text{ g of KCl}$		

From 1.00 g of  $\text{Cl}_2$  and excess K?

$\frac{1.00 \text{ g of Cl}_2}{70.906 \text{ g/mol}}$	=	0.0141 mol of $\text{Cl}_2$
$\frac{\text{Cl}_2}{1}$	=	$\frac{0.0141 \text{ mol of Cl}_2}{x}$
$\frac{\text{KCl}}{2}$	=	$\frac{x}{x}$
		$x = 0.0282 \text{ mol of KCl}$
$0.0282 \text{ mol} \times 74.551 \text{ g/mol} = 2.10 \text{ g of KCl}$		

# Chemistry Unit 9 Ws 2 Answers Stoichiometry

**DJ Losen**



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