Review Questions

1. 1   2. 3   3. 2
4. 4   5. 2   6. 2
7. 1   8. 3   9. 3

10. The plum padding model pictures the atom with an even distribution of positive charges, with electrons throughout the atom. The planetary model pictures the nucleus at the center of the atom with electrons in orbits.

11. Rutherford fired positively charged alpha particles at gold foil. He expected them to pass through with little deflection. When some were greatly deflected, he reasoned that they must have been repelled by a massive positively charged particle, which was called the nucleus.

12. The Rutherford model pictured the electrons in fixed, orbiting positions. The wave-mechanical model pictures electrons as possessing both wave and particle properties.

13. 3   14. 3   15. 1
16. 3   17. 4   18. 4
19. 2   20. 4   21. 2
22. 1   23. 1   24. 3
25. 3   26. 2   27. 2
28. 3   29. 3   30. 2
31. 3   32. 4   33. 4
34. 4   35. 1   36. 2
37. 1   38. 1   39. 3
40. 1   41. 2   42. 3
43. 2   44. 4   45. 4
46. 3   47. 3   48. 4
49. 2   50. 3   51. 2
52. 2   53. 3   54. 1
55. 3   56. 3

57. Na $1s^2 2s^2 2p^6 3s^1$
Mg $1s^2 2s^2 2p^6 3s^2$
Al $1s^2 2s^2 2p^6 3s^2 3p^1$
Si $1s^2 2s^2 2p^6 3s^2 3p^2$
P $1s^2 2s^2 2p^6 3s^2 3p^3$
S $1s^2 2s^2 2p^6 3s^2 3p^4$
Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Ar $1s^2 2s^2 2p^6 3s^2 3p^6$

58. 1   59. 4   60. 1
61. 1   62. 2   63. 4
64. 4   65. 2   66. 2
67. 2   68. 2   69. 1

70. The mixture contains at least three parts. One solid remained on the filter paper. The other two parts consist of a liquid with a solid dissolved in it.

71. If the calcium and sulfur can be separated by physical means, the substance is a mixture. If chemical means are required to separate them, it is a compound.

72. The picture should show several of each type of molecule evenly mixed.

73. A mixture of iron and oxygen would have the iron present as a solid and the oxygen present as a gas. The iron would be attracted to a magnet. A glowing splint would ignite into a flame if placed into the oxygen gas. If the iron and oxygen were combined as a compound, it would be a solid that is not attracted to a magnet.

74. Pepperoni pizza is heterogeneous because individual parts can be seen.

75. A is a mixture of an element and a compound; one molecule contains identical atoms and the other one contains different types of atoms. B is a compound; all particles are identical and are made up of two different types of atoms. C is an element; all molecules are identical and each is made up of two identical atoms. D is a mixture of an element and two different compounds; one molecule contains identical atoms and the other two contain different types of atoms, but they differ from each other.

76. Container D is homogeneous because the particles are mixed on the molecular level.

77. The two substances are different materials because their composition varies.

Regents Practice Questions

Part A

1. 1   2. 1   3. 1
4. 3   5. 1   6. 4
7. 2   8. 3   9. 3
10. 2   11. 3   12. 1
13. 2   14. 3

Part B

15. 4   16. 3   17. 1
18. 3   19. 1   20. 3
21. 1   22. 3   23. 2
24. 4

Part C

25. Because the two nuclei have different numbers of protons, they have different amounts of attraction for electrons. As a result, the amounts of energy separating the energy levels differ for the two elements.

26. The samples are not the same substance because the percent of each element differs from sample to sample.

27. Because he assumed the atom to have an even distribution of positive charges, he did not expect a concentration of positive charges that would strongly deflect or bounce back the alpha particles.

28. The spectra are based on the elements present in the compound. If some lines matched and some didn't, the compounds have a least one element in common and at least one differing element.

29. Column 1: 1, 2, 1; column 2: K, 39, 19; column 3: 92, 146, 92

30. $(0.9778)(12) + (0.0222)(13) = 12.02$

31. They contain different numbers of neutrons.

32. $(19.99 \times .909) + (20.99 \times 0.03) + (21.99 \times 0.88)$

33. 20

34. $\dfrac{(12.00 \times 9.93) + (13.00 \times 1.07)}{100}$

35. Both isotopes contain 6 protons but different numbers of neutrons. Or: An isotope with a mass of 12.00 contains 6 protons and 6 neutrons, while the isotope with a mass of 13.00 contains 6 protons and 7 neutrons.

[Chemquest 15 Answer](#)

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