

ChemQuest 31

Using Moles with Formulas

Name: _____

Date: _____

Hour: _____

Information: Percent Composition

Sometimes it is needful to know the composition of a compound. For example, 39.3% of the mass of sodium chloride is due to sodium. The other 60.7% of the mass is from chlorine. So, in a 100 g sample of sodium chloride, there are 39.3 g of sodium and 60.7 g of chlorine. This type of data is known as **percent composition**. The percent composition tells you the percentage by mass of an element in a compound. There is a convenient formula for finding the percent composition of an element in a compound:

$$\text{percent composition of element "X"} = \frac{\text{mass of X in one mole of the compound}}{\text{mass of one mole of the compound}} \cdot 100$$

(obtained from periodic table) →

(obtained from periodic table) →

Let us look at how the percent composition of calcium (Ca) in calcium chloride (CaCl_2) was determined.

$$\text{percent composition of Ca} = \frac{\text{mass of Ca in one mole of CaCl}_2}{\text{mass of one mole of CaCl}_2} \cdot 100$$

$$\text{percent composition of Ca} = \frac{40.1 \text{ g}}{111.1 \text{ g}} \cdot 100 = 36.1\%$$

← from periodic table for calcium

← from periodic table for calcium + 2 chlorines:
40.1 + 2(35.5) = 111.1

As another example, consider calculating the percent composition of nitrogen in Ca_3N_2 :

$$\text{percent composition of N} = \frac{28.0 \text{ g}}{148.3 \text{ g}} \cdot 100 = 18.9\%$$

← From periodic table for 2 nitrogen atoms: 2(14.0)=28.0

← from periodic table for 3 calcium + 2 nitrogen atoms
3(40.1) + 2(14.0) = 148.3

Critical Thinking Questions

- Verify that in C_4H_{10} the percent composition of carbon is approximately 82.6%.

$$\frac{4(12.0)}{[4(12.0) + 10(1.01)]} \cdot 100 = 82.6\%$$

Note: 12.0 and 1.01 are the molar masses for carbon and hydrogen on the periodic table.

Chemquest 31 Answers

Mike Jess



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Table of Contents Chemquest 31 Answers

1. Understanding the eBook Chemquest 31 Answers
 - The Rise of Digital Reading Chemquest 31 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemquest 31 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemquest 31 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 31 Answers
 - Personalized Recommendations

- Chemquest 31 Answers User Reviews and Ratings
- Chemquest 31 Answers and Bestseller Lists
- 5. Accessing Chemquest 31 Answers Free and Paid eBooks
 - Chemquest 31 Answers Public Domain eBooks
 - Chemquest 31 Answers eBook Subscription Services
 - Chemquest 31 Answers Budget-Friendly Options
- 6. Navigating Chemquest 31 Answers eBook Formats
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 - Chemquest 31 Answers Compatibility with Devices
 - Chemquest 31 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Chemquest 31 Answers
 - Interactive Elements Chemquest 31 Answers
- 8. Staying Engaged with Chemquest 31 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemquest 31 Answers
- 9. Balancing eBooks and Physical Books Chemquest 31 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemquest 31 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemquest 31 Answers
 - Setting Reading Goals Chemquest 31 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemquest 31 Answers
 - Fact-Checking eBook Content of Chemquest 31 Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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