

## ChemQuest 27

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Name: \_\_\_\_\_

Date: \_\_\_\_\_

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# Covalent Bonding

### Information: Terminology

Recall that an ionic bond results from the combination of a metal and a nonmetal. (Note: hydrogen is a nonmetal even though it is on the left side of the periodic table.)

A covalent bond is the type of bond between two nonmetals. Covalent bonds are formed by neutral atoms that share electrons rather than by charged ions. When a compound is formed by sharing electrons, the compound is called a molecule or molecular compound. It is important to note that ionic compounds are not called molecules. The largest class of molecules is called organic molecules. Carbon is the distinguishing mark of organic compounds.

### Critical Thinking Questions

- Use your knowledge of metals and nonmetals to identify each of the following as having ionic (I) or covalent (C) bonds.  
a)  $\text{CaCl}_2$       b)  $\text{MgO}$       c)  $\text{CH}_4$       d)  $\text{N}_2\text{O}_5$       e)  $\text{Al}(\text{NO}_3)_3$       f)  $\text{SO}_2$
- The compound  $\text{NH}_4\text{NO}_3$  is an ionic compound even though there isn't a metal in the compound. Explain why? (Hint: Think about what  $\text{NH}_4\text{NO}_3$  is made of.)
- Circle any of the following compounds that would properly be called a "molecule".  
a)  $\text{H}_2\text{O}$       b)  $\text{CCl}_4$       c)  $\text{NaCl}$       d)  $\text{Mg}_3\text{P}_2$       e)  $\text{N}_2\text{O}_5$
- Which compound from question 3 would be classified as "organic"? Why?

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**G Orfield**



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