

## Assessment

## Chapter Test B

## Chapter: Chemical Formulas and Chemical Compounds

**PART I** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Changing a subscript in a correctly written chemical formula will
- change the electron configuration of that element.
  - change the charges on the other ions in the compound.
  - change the formula so that it no longer represents the same compound.
  - have no effect on the formula.
- \_\_\_\_\_ 2. Using the Stock system of nomenclature,  $\text{Cr}_2(\text{SO}_4)_3$  is named
- chromium(II) sulfate.
  - chromic sulfate.
  - dichromium trisulfate.
  - chromium(III) sulfate.
- \_\_\_\_\_ 3. In a polyatomic ion, the algebraic sum of the oxidation numbers of all atoms is equal to
- 0.
  - 10.
  - the number of atoms in the ion.
  - the charge on the ion.
- \_\_\_\_\_ 4. The first part of the name of a binary ionic compound is the
- cation.
  - polyatomic ion.
  - oxyanion.
  - anion.
- \_\_\_\_\_ 5. The empirical formula may not represent the actual composition of a(n)
- ionic compound.
  - crystal.
  - salt.
  - molecular compound.
- \_\_\_\_\_ 6. The formula for phosphoric acid is
- $\text{HPO}_4$
  - $\text{H}(\text{PO}_4)_2$
  - $\text{H}_2\text{PO}_3$
  - $\text{H}_3\text{PO}_4$

# Chapter Chemical Formulas And Compounds Test

**R Barnett**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle overlapping the right side of the bar.

## **Chapter Chemical Formulas And Compounds Test:**

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## **Table of Contents Chapter Chemical Formulas And Compounds Test**

1. Understanding the eBook Chapter Chemical Formulas And Compounds Test
  - The Rise of Digital Reading Chapter Chemical Formulas And Compounds Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Chemical Formulas And Compounds Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter Chemical Formulas And Compounds Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Chemical Formulas And Compounds Test
  - Personalized Recommendations
  - Chapter Chemical Formulas And Compounds Test User Reviews and Ratings
  - Chapter Chemical Formulas And Compounds Test and Bestseller Lists

5. Accessing Chapter Chemical Formulas And Compounds Test Free and Paid eBooks
  - Chapter Chemical Formulas And Compounds Test Public Domain eBooks
  - Chapter Chemical Formulas And Compounds Test eBook Subscription Services
  - Chapter Chemical Formulas And Compounds Test Budget-Friendly Options
6. Navigating Chapter Chemical Formulas And Compounds Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter Chemical Formulas And Compounds Test Compatibility with Devices
  - Chapter Chemical Formulas And Compounds Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter Chemical Formulas And Compounds Test
  - Highlighting and Note-Taking Chapter Chemical Formulas And Compounds Test
  - Interactive Elements Chapter Chemical Formulas And Compounds Test
8. Staying Engaged with Chapter Chemical Formulas And Compounds Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter Chemical Formulas And Compounds Test
9. Balancing eBooks and Physical Books Chapter Chemical Formulas And Compounds Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter Chemical Formulas And Compounds Test
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter Chemical Formulas And Compounds Test
  - Setting Reading Goals Chapter Chemical Formulas And Compounds Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter Chemical Formulas And Compounds Test
  - Fact-Checking eBook Content of Chapter Chemical Formulas And Compounds Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- 
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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