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Teacher's resource answers

Chapter 21

1 $K_a = 10^{-7.5} = 10^{-7.5}$ [1]

$= 6.31 \times 10^{-8} \text{ mol dm}^{-3}$ [2]

(take 1 mark off if units are incorrect)

$[\text{H}^+] \sqrt{K_a \times [\text{H}_2\text{PO}_4^-]} = 7.94 \times 10^{-5} \text{ mol dm}^{-3}$

(this is because $K_a = \frac{[\text{H}^+][\text{HPO}_4^{2-}]}{[\text{H}_2\text{PO}_4^-]}$) [1]

$\text{pH} = -\log_{10}[\text{H}^+] = 4.10$ [1]

2 a A buffer is a mixture of a weak acid and its conjugate base. [1]

It resists changes in pH. [1]

b The equilibrium present in the buffer is $\text{H}_2\text{PO}_4^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}^+$ [1]

If unwanted H^+ ions are added they combine with the HPO_4^{2-} [1]

to form the largely undissociated weak acid H_2PO_4^- and are therefore removed, leaving the pH near constant. [1]

If unwanted OH^- ions are added they are neutralised by the H^+ ions thus disturbing the equilibrium. [1]

More H_2PO_4^- dissociates until the K_a returns to its proper value and pH remains near constant. [1]

c $K_a = \frac{[\text{H}^+][\text{HPO}_4^{2-}]}{[\text{H}_2\text{PO}_4^-]}$ [1]

rearranging the formula gives

$[\text{H}^+] = K_a \times \frac{[\text{H}_2\text{PO}_4^-]}{[\text{HPO}_4^{2-}]}$ [1]

$[\text{H}^+] = 6.31 \times 10^{-8} \times \frac{0.01}{0.06} = 1.05 \times 10^{-8} \text{ mol dm}^{-3}$ [1]

$\text{pH} = -\log_{10}[\text{H}^+] = 7.98$ [1]

d The equilibrium $\text{H}_2\text{PO}_4^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}^+$ would be displaced to the left [1]

in order to remove the added H_2PO_4^- ions (Le Chatelier's principle). [1]

Therefore, $[\text{H}^+]$ increases and pH decreases. [1]

3 a Ca^{2+} and PO_4^{3-} [2]

b $\text{Ca}_3(\text{PO}_4)_2 \rightleftharpoons 3\text{Ca}^{2+} + 2\text{PO}_4^{3-}$ [1]

c $K_{sp} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2$ [1]

d solubility $= 2.48 \times 10^{-8} \text{ mol dm}^{-3}$ [1]

$[\text{Ca}^{2+}] = 3 \times 2.48 \times 10^{-8} = 7.44 \times 10^{-8} \text{ mol dm}^{-3}$ [1]

and $[\text{PO}_4^{3-}] = 2 \times 2.48 \times 10^{-8} = 4.96 \times 10^{-8} \text{ mol dm}^{-3}$ [1]

$K_{sp} = (7.44 \times 10^{-8})^3 \times (4.96 \times 10^{-8})^2 = 1.01 \times 10^{-28} \text{ mol}^5 \text{ dm}^{-15}$ [2]

(take one mark off for incorrect units)

e There is a strong attraction between the +2 calcium ions and the -3 phosphate ions. [1]

Therefore, when the gaseous ions form the ionic lattice a lot of energy is produced as strong ionic bonds are formed. [1]

f The enthalpies of hydration of the Ca^{2+} ions and the PO_4^{3-} ions are not sufficiently negative (enough) [1]

to overcome the large amount of energy needed to produce the gaseous ions from the ionic lattice (-L.E.). [1]

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Table of Contents Chapter Resource Answers

1. Understanding the eBook Chapter Resource Answers
 - The Rise of Digital Reading Chapter Resource Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Resource Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Resource Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Resource Answers
 - Personalized Recommendations
 - Chapter Resource Answers User Reviews and Ratings
 - Chapter Resource Answers and Bestseller Lists
5. Accessing Chapter Resource Answers Free and Paid eBooks
 - Chapter Resource Answers Public Domain eBooks
 - Chapter Resource Answers eBook Subscription Services
 - Chapter Resource Answers Budget-Friendly Options
6. Navigating Chapter Resource Answers eBook Formats

- ePub, PDF, MOBI, and More
- Chapter Resource Answers Compatibility with Devices
- Chapter Resource Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Resource Answers
 - Highlighting and Note-Taking Chapter Resource Answers
 - Interactive Elements Chapter Resource Answers
- 8. Staying Engaged with Chapter Resource Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Resource Answers
- 9. Balancing eBooks and Physical Books Chapter Resource Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Resource Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Resource Answers
 - Setting Reading Goals Chapter Resource Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Resource Answers
 - Fact-Checking eBook Content of Chapter Resource Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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