

Periodic Trends

Can the properties of an element be predicted using a periodic table?

Why?

The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

3 1. Consider the data in Model 1 on the following page.

a. Each element has three numbers listed under it. Which value represents the atomic radius?

The first number below the diagram

b. What are the units for the atomic radius?

pm picometers 10^{-12} m

c. Write a complete sentence to convey your understanding of atomic radius. *Note:* You may not use the word "radius" in your definition.

Distance from center of atom (nucleus) to outer electrons

2 2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.

Increases $H < Li < Na < K$ $F < Cl < Br$
 $C < Si < Ge$

3 3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. *Hint:* You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.

As the number of electrons increases, new energy levels need to be added further from the nucleus, increasing the radius.

2 4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods. *< exception*

Decreases $Li > Be > B > C > N > O > F > Ne$
 $Na > Mg > Al > Si > P > S > Cl > Ar$

5 5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.

Highest energy level is constant, but the number of protons in the nucleus increases, increasing the charge of the nucleus and the attractive force on the electrons, pulling the highest energy level closer to the nucleus.

Chemistry Periodic Trends Answer Key

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1. Understanding the eBook Chemistry Periodic Trends Answer Key
 - The Rise of Digital Reading Chemistry Periodic Trends Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Periodic Trends Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Periodic Trends Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Periodic Trends Answer Key
 - Personalized Recommendations
 - Chemistry Periodic Trends Answer Key User Reviews and Ratings
 - Chemistry Periodic Trends Answer Key and Bestseller Lists
5. Accessing Chemistry Periodic Trends Answer Key Free and Paid eBooks
 - Chemistry Periodic Trends Answer Key Public Domain eBooks
 - Chemistry Periodic Trends Answer Key eBook Subscription Services
 - Chemistry Periodic Trends Answer Key Budget-Friendly Options
6. Navigating Chemistry Periodic Trends Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry Periodic Trends Answer Key Compatibility with Devices
 - Chemistry Periodic Trends Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Periodic Trends Answer Key
 - Highlighting and Note-Taking Chemistry Periodic Trends Answer Key
 - Interactive Elements Chemistry Periodic Trends Answer Key
8. Staying Engaged with Chemistry Periodic Trends Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Periodic Trends Answer Key
9. Balancing eBooks and Physical Books Chemistry Periodic Trends Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Periodic Trends Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Periodic Trends Answer Key
- Setting Reading Goals Chemistry Periodic Trends Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Periodic Trends Answer Key
- Fact-Checking eBook Content of Chemistry Periodic Trends Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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