

## Acid Base Practice Questions

- 1) Write equations for the dissociation of the following compounds in aqueous solution. Identify the acid base conjugate pairs in each case (use appropriate arrows for strong and weak species)
- $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \leftrightarrow \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$
  - $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \leftrightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$
  - $\text{NH}_3 + \text{H}_2\text{O} \leftrightarrow \text{NH}_4^+ + \text{OH}^-$
  - $\text{HNO}_3 + \text{H}_2\text{O} \leftrightarrow \text{NO}_3^- + \text{H}_3\text{O}^+$
- 2) Define the term amphiprotic. Write appropriate equations to show the amphiprotic nature of water.  
Can accept or donate a proton
- $\text{NH}_3 + \text{H}_2\text{O} \leftrightarrow \text{NH}_4^+ + \text{OH}^-$  (acting as an acid)
  - $\text{HNO}_3 + \text{H}_2\text{O} \leftrightarrow \text{NO}_3^- + \text{H}_3\text{O}^+$  (acting as a base)
- 3) Calculate the pH of the following solutions
- $[\text{H}^+] = 2.5 \times 10^{-5} \text{ mol/L}$  4.6
  - $[\text{OH}^-] = 1.8 \times 10^{-4} \text{ mol/L}$  10.25
  - 0.006 M NaOH 11.77
- 4) Compare the  $[\text{H}^+]$  in solutions with the pH of 3 and 5 (pH 3 has 100x higher  $[\text{H}^+]$ )
- 5) Calculate the concentration of  $\text{H}^+$  and  $\text{OH}^-$  ions in 0.005 M solution of NaOH.
- $[\text{OH}^-] = 0.005$
  - $[\text{H}^+] = 2.0 \times 10^{-12}$
- 6) Calculate the pH of a 0.0004 M solution of  $\text{Ba}(\text{OH})_2$  10.9
- 7) Describe how you would experimentally distinguish a strong base and a strong acid.
- Test pH
  - React with a metal
- 8) Consider the reaction:  $2\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$
- What volume of 2.0 M ammonia is needed to neutralize 50 mL of a 0.5 M sulfuric acid? 25mL
  - Is the ammonium sulphate an acidic, basic or a neutral salt? Acidic
- 9) What are indicators? What factors must be considered when selecting an indicator for a titration?
- Substances that change colour over a specific pH range (called endpoint of indicator)
  - The pH at equivalence for titration must fall within the endpoint of indicator
- 10) Calculate the pH of (pH of a weak acid)
- 0.005 M solution of ethanoic acid ( $K_a = 1.76 \times 10^{-5}$ ) 3.52
  - 0.003 M solution of ammonia ( $K_b = 1.8 \times 10^{-5}$ ) 10.4
- 11) Calculate the
- pKa for acetic acid ( $K_a = 1.76 \times 10^{-5}$ ) 4.75
  - pKb for  $\text{NH}_3$  ( $K_b = 1.8 \times 10^{-5}$ ) 4.74
- 12) How are  $K_a$ , pKa and  $K_b$  and pKb related? (higher  $K_a$ , lower pKa) (higher  $K_b$ , lower pKb)
- 13) What are buffers?
- Substances that resist change in pH when small amounts of a strong base/acid is added to them
- 14) What are the main components of an acidic buffer, a basic buffer? Give examples.
- Acidic buffer: weak acid and a salt containing its conjugate base
  - Basic Buffer: weak base and a salt containing its conjugate acid

# Acid Base Equilibrium Practice Problems Multiple Choice

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