Summary: Acid-Base Reactions in Organic Chemistry

- Acid-base reactions involve transfers of a proton (H^+) between an **acid** and a **base**
- Every acid-base reaction has **4** components (**acid, base, conjugate acid, conjugate base**)

Form	Break
$H–NH_2$	H–F

- The **conjugate acid** is the product formed through protonation of the **base**
- The **conjugate base** is the product that remains after the **acid** has transferred its proton
- The strength of an acid is directly related to the stability of its conjugate base.

"The stronger the acid, the weaker the conjugate base"
not "the stronger the acid, the weaker the base"

Example: across a row of the periodic table, acidity increases with electronegativity of the atom, which is better able to stabilize negative charge (C < N < O < F)

least stable conjugate base $:CH_3^- < NH_2^- < OH^- < F^-$ most stable (weakest) conjugate base

- Acid base equilibrium favors the direction in which a **stronger acid and stronger base** combine to give a **weaker acid and weaker base.**

Form	Break
$H–NH_2$	H–F

- H–F is a stronger acid than $H–NH_2$ (the conjugate acid),
- NH_2^- is a stronger base than F^- (the conjugate base)

So the reaction equilibrium will favor formation of the weaker acid (NH_3) and weaker base (F-)

Acid Base Reactions Answers

G Thomas

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