

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~Sulfuric acid~~ **H<sub>2</sub>SO<sub>4</sub>**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~potassium hydroxide~~ **KOH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
HClO<sub>4</sub> is STRONG with a lower cell completely dissociate in water

- HCl**      ~~HCl~~      ~~HNO<sub>3</sub>~~      ~~Mg(OH)<sub>2</sub>~~      **HNO<sub>2</sub>**      **Ca(OH)<sub>2</sub>**      ~~NaOH~~      **HClO<sub>4</sub>**

3. What is the formula for hydronium? **H<sub>3</sub>O<sup>+</sup>**

4.  $\text{H}_2\text{C}_2\text{O}_4$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for the reaction.  **$\text{H}_2\text{C}_2\text{O}_4 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{C}_2\text{O}_4^+ + \text{HC}_2\text{O}_4^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is a weak acid.**
- Would you define  $\text{H}_2\text{C}_2\text{O}_4$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{H}_2\text{C}_2\text{O}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{H}_2\text{C}_2\text{O}_4$ .  **$\text{HC}_2\text{O}_4^-$**
- Identify  $\text{H}_2\text{C}_2\text{O}_4$  as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5.  $\text{F}^-$  will react with water and gain a  $\text{H}^+$  ion in the process.

- Write a balanced equation for the reaction.  **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{F}^-$  is a weak base.**
- Would you define  $\text{F}^-$  as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{F}^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of  $\text{F}^-$ . **HF**
- Identify  $\text{F}^-$  as polyprotic, amphoteric, or neither. **neither**

6.  $\text{HCO}_3^-$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for the reaction.  **$\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3^* + \text{CO}_3^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{HCO}_3^-$  is a weak acid.**
- Would you define  $\text{HCO}_3^-$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{HCO}_3^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{HCO}_3^-$ .  **$\text{CO}_3^{2-}$**
- Identify  $\text{HCO}_3^-$  as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only one H to be a donor) and a H to give (one can give an acid).**

# Acids And Bases Answer Key

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