

CHAPTER 6

STUDY GUIDE

Ionic Compounds and Metals

Lesson 6.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

| | | | | |
|---------------|-----------|-----------------------------|---------|-------------|
| chemical bond | electrons | energy level | ions | noble gases |
| nucleus | octet | pseudo-noble gas formations | valence | |

The force that holds two atoms together is called a(n) (1) Chemical Bond.

Such an attachment may form by the attraction of the positively charged

(2) Nucleus of one atom for the negatively charged

(3) Electron of another atom, or by the attraction of charged atoms, which

are called (4) Ions. The attractions may also involve

(5) Valence electrons, which are the electrons in the outermost

(6) Energy level. The (7) Noble Gas are a family of

elements that have very little tendency to react. Most of these elements have a set of eight outermost

electrons, which is called a stable (8) Octet. The relatively stable electron

structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9)

Pseudo-Noble gas formations.

For each statement below, write *true* or *false*.

- False 10. A positively charged ion is called an anion.
- True 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- False 12. Elements tend to react so that they acquire the electron structure of a halogen.
- True 13. A sodium atom tends to lose one electron when it reacts.
- True 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- False 15. A Cl^- ion is an example of a cation.
- True 16. The ending *-ide* is used to designate an anion.
- False 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

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RJ Shavelson



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