Name: KEY Block: ____

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg_2O
3. gold (III) chloride $AuCl_3$
4. thallium (III) bromide $TlBr_3$
5. bismuth (V) oxide Bi_2O_5
6. terbium (III) oxide Tb_2O_3
7. uranium (VI) oxide UO_3
8. protactinium (IV) bromide $PaBr_4$
9. cerium (III) oxide Ce_2O_3
10. arsenic (V) sulphide As_2S_5
11. manganese (III) oxide Mn_2O_3
12. vanadium (II) bromide VBr_2
13. niobium(V) oxide Nb_2O_5
14. titanium (III) oxide Ti_2O_3
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co_3P_2
18. tin (II) oxide SnO
19. thulium (II) bromide $TmBr_2$
20. copper (II) bromide $CuBr_2$

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. $SnCl_4$ Tin (IV) Chloride
2. $BiBr_5$ Bismuth (V) Bromide
3. PoO_2 Polonium (IV) Oxide
4. PbI_2 Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au_2O_3 Gold (III) Oxide
8. $FeCl_2$ Iron (II) Chloride
9. PdF_4 Palladium (IV) Fluoride
10. Os_2O_3 Osmium (III) Oxide
11. $MoBr_2$ Molybdenum (II) Bromide
12. VCl_5 Vanadium (V) Chloride
13. Mn_2O_3 Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np_2O_3 Neptunium (III) Oxide
16. V_2O_5 Vanadium (V) Oxide

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