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9 CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

   Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

   The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

   The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

   a. alkali metal

   b. alkaline earth metal

   c. transition metal

   d. nonmetal

5. The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the lesser charge, and the suffix *-ic* to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

   The Stock system gives the actual charge of the ion.

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Chapter 9 Chemical Names And Formulas Answer Key

Didier Musso

Chapter 9 Chemical Names And Formulas Answer Key:

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