

## Chapter 7: Chemical Formulas and Chemical Compounds

### I. CHEMICAL NAMES AND FORMULAS

#### A. Significance of a Chemical Formula

1. A chemical formula shows the relative number of atoms of each kind in a chemical compound, with subscripts indicating the number of atoms of one type of element.
2. In a molecular formula, the chemical formula shows the number of atoms of each element in one molecule of the compound.
  - a) For example,  $H_2O$ , the chemical formula for water, indicates the presence of 2 hydrogen atoms and 1 oxygen atom.
3. In an ionic compound, the chemical formula represents one formula unit, which is the simplest ratio of the compound, positive ions (cations) and negative ions (anions).

#### B. Monatomic Ions

1. Many main-group elements can become ions by gaining or losing electrons to have noble-gas configurations.
2. Monatomic ions are ions formed from a single atom.
  - a) Group 1 metals lose one electron; Group 2 elements lose two electrons. They will have +1 and +2 charges, respectively.
  - b) The nonmetals in Group 15, 16, and 17 gain electrons and form anions.
  - c) Metals with *d* orbitals are slightly different, as they can form +2, +3, or even +1 and +4. Some of them can even form ions of different charges.
    - i) Refer to the chart on page 221 of Modern Chemistry published by Holt, Rinehart, and Winston.

#### C. Naming Monatomic Ions

1. Monatomic cations are identified by their normal name.
2. Monatomic anions, however, have their ending dropped and the ending, -ide is added to the root name.

#### D. Binary Ionic Compounds

1. Binary compounds are compounds composed of two elements.
2. In a binary ionic compound, the total numbers of positive charges and negative charges must be equal.

# Chapter Chemical Formulas And Compounds

**Thomas Griffiths**



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