

### Open-Ended 13

## Electron Configuration

Name \_\_\_\_\_

Date \_\_\_\_\_

Year \_\_\_\_\_

### Introduction: Energy of Sublevels

Sub-levels have different energies of energy. For example, orbitals in the  $3s$  sublevel have more energy than orbitals in the  $2p$  sublevel. The following activities will help you learn how to write energy levels.



Circle the  $s$  orbitals and underline the  $p$  orbitals. Which orbitals are higher in energy?



Now list the  $s$  orbitals again, this time in order of increasing energy. Do you agree with the order of energy that you determined for  $d$  orbitals in a higher energy level than  $s$  orbitals?

When drawing energy levels, be sure to include the correct number of orbitals. (This is called the **Aufbau Principle**). An orbital will only be filled with electrons if it has a lower energy than the orbitals that immediately after it is already filled. Additionally, no orbital can be filled with more than two electrons. (This is called the **Pauli Exclusion Principle**). When two electrons occupy the same orbital they must spin in opposite directions—represented by the up and down arrows.

### Critical Thinking Questions

- a. How many electrons would you need to determine how high filling the  $d$  orbitals? **Indicating how filled all the orbitals you filled shows electron configuration for  $3d$ .**

b. Why do the  $d$  orbitals have a higher energy than the  $s$  orbitals? **Why not?  $d$  orbitals are higher in energy than  $s$  orbitals. The  $d$  orbitals are higher in energy.**
- a. How many electrons would you need to determine how high filling the  $f$  orbitals? **Indicating how filled all the orbitals you filled for  $4f$ — $4s$ — $4p$ — $4d$ .**

b. Why do the  $f$  orbitals have a higher energy than the  $d$  orbitals? **The  $f$  orbitals are higher in energy than  $d$  orbitals.**

### Introduction: Hund's Rule

Electrons prefer to be "up" or "spinned". Electrons in the same orbital will be paired up. Paired electrons are in the same orbital. For example, two electrons in the same orbital will be paired up.



If available, electrons in the  $s$  orbital before the  $d$  orbitals. The  $d$  orbitals are higher in energy than the  $s$  orbitals. The  $d$  orbitals are higher in energy than the  $s$  orbitals.

### Choice A or Choice B



How's rule is to fill orbitals in order of increasing energy. If the orbitals are in the same energy level, the orbitals will be filled with electrons. The orbitals will be filled with electrons. The orbitals will be filled with electrons.

The following are the electron configurations for the two elements, nitrogen and oxygen. Write the energy level diagram for each element. How do the  $d$  orbitals compare to the  $s$  orbitals?



### Critical Thinking Questions

- a. How many "up" or "spinned" electrons are in a nitrogen atom? **1**
- a. Why does the  $d$  orbitals have a higher energy than the  $s$  orbitals? Why not? **Why not?  $d$  orbitals are higher in energy than  $s$  orbitals.**
- Why do the  $d$  orbitals have a higher energy than the  $s$  orbitals? **The  $d$  orbitals are higher in energy than the  $s$  orbitals.**
- Why do the  $d$  orbitals have a higher energy than the  $s$  orbitals? **The  $d$  orbitals are higher in energy than the  $s$  orbitals.**

# Chemquest 18 Answer Key

**Brendan G. Carr**

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