

**Worksheet 2: Binary Ionic Compounds (variable charge metals) - metals from groups 3 - 12 and 14 or 4A (Pb and Sn) - charge of anion is constant - use charge of anion to figure out charge of metal**

1. Name the following ionic compounds:

a. $\text{SnO}_2$	oxide = $\text{O}^{-2}$	Tin (IV) oxide
b. $\text{Mn}_2\text{O}_7$	oxide = $\text{O}^{-2}$	Manganese (VII) oxide
c. $\text{FeN}$	nitride = $\text{N}^{-3}$	Iron (III) nitride
d. $\text{Cu}_3\text{N}_2$	nitride = $\text{N}^{-3}$	Copper (II) nitride
e. $\text{TiF}_3$	fluoride = $\text{F}^{-1}$	Titanium (III) fluoride
f. $\text{Cu}_2\text{S}$	sulfide = $\text{S}^{-2}$	Copper (I) sulfide
g. $\text{Fe}_2\text{S}_3$	sulfide = $\text{S}^{-2}$	Iron (III) sulfide
h. $\text{CuBr}$	bromide = $\text{Br}^{-1}$	Copper (I) bromide
i. $\text{Co}_3\text{N}_2$	nitride = $\text{N}^{-3}$	Cobalt (II) nitride
j. $\text{CoF}_2$	fluoride = $\text{F}^{-1}$	Cobalt (II) fluoride

2. Write the formula for the following ionic compounds:

a. Copper (I) nitride	$\text{Cu}_3\text{N}$
b. Cobalt (I) fluoride	$\text{CoF}$
c. Titanium (IV) oxide	$\text{TiO}_2$
d. Iron (II) nitride	$\text{Fe}_3\text{N}_2$
e. Iron (III) chloride	$\text{FeCl}_3$
f. Copper (II) oxide	$\text{CuO}$
g. Rhodium (II) oxide	$\text{RhO}$
h. Tin (IV) sulfide	$\text{SnS}_2$
i. Manganese (IV) nitride	$\text{Mn}_3\text{N}_4$
j. Copper (I) fluoride	$\text{CuF}$
k. Cobalt (II) sulfide	$\text{CoS}$
l. Iron (III) oxide	$\text{Fe}_3\text{O}_2$

# Chemical Formulas Answer Key

**Katrin Zwirgmaier**



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