

Chemical Names and Formulas

❖ Key Vocabulary

Monoatomic Ion – Ions formed from a single atom

Binary compound – Compounds composed of two elements

Nomenclature – the “naming system” of binary ionic compounds

Oxyanion – polyatomic ions that contain oxygen

Salt – an ionic compound composed of a cation and the anion from an acid

Chemical Formulas

Subscripts after an elemental symbol represent the number of atoms of the element in the compound

Molecular formulas represent how many atoms of each element are in a compound.
(C_2H_6 , H_2O , $C_6H_{12}O_6$)

Ionic formulas show the simplest ratio, cations to anions, of atoms in the lattice

Parenthesis are used around a polyatomic ion in a chemical formula

Examples $Al_2(SO_4)_3$, $Cu(PO_4)_2$, $Sr(NO_3)_2$

Monoatomic Ions

– Ions formed from a single atom

Group 1 elements may lose one electron = 1+

Group 2 elements may lose two electrons = 2+

Group 13 elements may lose three electrons = 3+

Group 17 elements may gain one electron = 1-

Group 16 elements may gain two electrons = 2-

Group 15 elements may gain three electrons = 3-

Transition metals have various possibilities (no pattern)

Naming monoatomic Ions

Monoatomic cations are simply called by the element's name

*the stock system is used when there is more than one possibility for an element (more on this later)

Monoatomic anions end in IDE such as Fluorine – Fluoride, Oxygen – Oxide

(reference table 1 on page 221)

Binary Ionic Compounds

1. Write the symbols for the ions with the cation first ($Al^{3+} O^{2-}$)

2. Cross over the charges writing the charge of one ion as the subscript for the other ion ($Al_2^{3+} O_3^{2-}$)

3. Multiply subscript and superscript for each ion then find the lowest whole number ratio (Al_2O_3)

Name it!!! Aluminum Oxide

Examples: $Zn^{2+} I^-$ $Zn^{2+} S^{2-}$

(More on p 223 practice problems)

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Puja Mehta



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