

TOPIC 5: AMOUNT OF SUBSTANCE

CP2 – FIND THE CONCENTRATION OF A SOLUTION OF SODIUM HYDROXIDE

Assessment criteria:

	1a		1b		1c	
CP2	Correctly uses appropriate instrumentation, apparatus and materials (including RT) to carry out investigative activities, experimental techniques and procedures with minimal assistance or prompting.		Identifies hazards and assesses risks associated with these hazards, making safety adjustments as necessary, when carrying out experimental techniques and procedures in the lab or field.		Makes accurate observations relevant to the experimental or investigative procedure.	
When is this assessed?	Usually during the experiment.		Established in risk assessment and during experiment.		Established in results table and usually during the experiment.	
Working towards		Student uses pipette and burette effectively, but requires some assistance, eg. reminder to wash or rinse an bubble, or to remove foam from the burette.		Student uses titrants, but incorrectly or inaccurately applies one titrant i.e. reads the buret for the wrong concentration.		Student obtains an accurate end point, with some assistance (titre and concordant end points).
		Student inaccurately makes a standard solution, with minimal assistance. Transferring their stock solution accurately, uses weighings, leaves the volumetric flask and does not check the volume. They do overshoot the volume. They then make the solution again accurately.		Students identify the hazards of 'alkaline acid' and 'alkaline hydroxide solution', but may not identify how to minimise risk.		Student makes accurate measurements, such as recording the volume on burette and mass of sodium acid, with some assistance.
		Student inaccurately reads burette and notes recording how to do this. Or, student does not correctly identify the end point of the reaction.		A procedure error identified, but one that's not basic for it level.		Student completes a final analysis using an indicator, but does not obtain concordant results.
Competent		Correct use of pipette and burette, ensuring that they are rinsed appropriately and that there are no air bubbles. Only one reminder required for foam.		Student uses titrants, prepared effectively to identify hazards and risks.		Student obtains an accurate end point (titre and concordant end points).
		Student accurately makes a standard solution. Transferring their stock solution accurately, uses weighings, leaves the volumetric flask and does not overshoot the volume.		Students identify the hazards of 'alkaline acid' and 'alkaline hydroxide solution'.		Student makes accurate measurements, such as recording the volume on burette and mass of sodium acid.
		Accurately reads the volume on the burette and notes the correct colour of the indicator for the end point of the reaction.		Students identify procedure of hazards and risk eg. those relating to glass breakage.		Student obtains concordant results.

PRE-LAB WORK:

You must write a risk assessment for the experiment. Below is an exemplar risk assessment for CP1, to give you guidance on how yours must be structured.

Following this, you must complete the risk assessment that is after the practical instructions, using the hazards provided.

Also, prior to the lab, make sure that you know how to do the following:

- ✓ That you know how to carry out a basic titration.
- ✓ Know the colours of methyl orange in acidic, neutral and alkaline solutions.
- ✓ That you know how to make a standard solution and why chemists make standard solutions.
- ✓ Know how to draw a results table and the degree of accuracy to which this can be recorded for a titration.

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