

# BALANCING REDOX REACTIONS

STEP BY STEP IN TWO WAYS...



## BOTH START THE SAME WAY

1. ASSIGN OXIDATION NUMBERS TO EACH ATOM



NOTE: OXIDATION NUMBERS ALWAYS SUM UP TO ZERO (NEUTRAL)

EXCEPT FOR IONS (IN THIS CASE, THE SULFATE ION)

3. BALANCE EACH OF THE ELEMENTS SEPARATELY

2. WRITE YOUR EQUATION AS TWO HALF REACTIONS (BALANCING ALL OTHERS BEFORE EITHER)

NOTE: ONLY O<sub>2</sub> & H<sub>2</sub> CAN BE ADDED TO EITHER SIDE OF EITHER REACTION



## HALF-REACTION METHOD (in acidic conditions)

3. BALANCE O WITH H<sub>2</sub>O



4. BALANCE H WITH H<sup>+</sup>



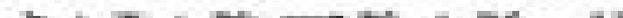
5. BALANCE CHARGES WITH ELECTRONS

NOTE: ALWAYS INCLUDE BOTH IN YOUR HALF REACTIONS AND CHECK THE OTHER



6. MAKE SURE THE NUMBER OF ELECTRONS IS THE SAME

NOTE: IN THIS CASE, BOTH ARE 2E, SO NO CHANGES TO EITHER HALF REACTION



7. PUT BOTH REACTIONS BACK TOGETHER



8. CHECK FOR ANYTHING THAT IS BOTH A REACTANT AND PRODUCT

NOTE: NONE IDENTIFIED, BUT BOTH H<sup>+</sup> AND H<sub>2</sub>O CAN BE



## BUT WHAT ABOUT BASIC CONDITIONS?

WHEN WE SAY "ACIDIC" OR "BASIC" WE ARE USING THE APPROXIMATE DEFINITION:

ACIDIC CONDITIONS: H<sup>+</sup> OR H<sub>2</sub>O CAN BE ADDED TO EITHER SIDE OF EITHER

BASIC CONDITIONS: OH<sup>-</sup> OR H<sub>2</sub>O CAN BE ADDED TO EITHER SIDE OF EITHER

8. USE THE NUMBER OF ELECTRONS TO THE NUMBER OF O<sub>2</sub> OR H<sub>2</sub> MOLECULES

NOTE: ONLY O<sub>2</sub> OR H<sub>2</sub> CAN BE ADDED TO THE SAME SIDE THAT YOU WANT



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**PT Brinkman**



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## **Table of Contents Balancing Redox Equations Study Guide Pg 117**

1. Understanding the eBook Balancing Redox Equations Study Guide Pg 117
  - The Rise of Digital Reading Balancing Redox Equations Study Guide Pg 117
  - Advantages of eBooks Over Traditional Books
2. Identifying Balancing Redox Equations Study Guide Pg 117
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Balancing Redox Equations Study Guide Pg 117
  - User-Friendly Interface
4. Exploring eBook Recommendations from Balancing Redox Equations Study Guide Pg 117
  - Personalized Recommendations
  - Balancing Redox Equations Study Guide Pg 117 User Reviews and Ratings
  - Balancing Redox Equations Study Guide Pg 117 and Bestseller Lists
5. Accessing Balancing Redox Equations Study Guide Pg 117 Free and Paid eBooks
  - Balancing Redox Equations Study Guide Pg 117 Public Domain eBooks
  - Balancing Redox Equations Study Guide Pg 117 eBook Subscription Services
  - Balancing Redox Equations Study Guide Pg 117 Budget-Friendly Options
6. Navigating Balancing Redox Equations Study Guide Pg 117 eBook Formats

- ePub, PDF, MOBI, and More
  - Balancing Redox Equations Study Guide Pg 117 Compatibility with Devices
  - Balancing Redox Equations Study Guide Pg 117 Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Balancing Redox Equations Study Guide Pg 117
    - Highlighting and Note-Taking Balancing Redox Equations Study Guide Pg 117
    - Interactive Elements Balancing Redox Equations Study Guide Pg 117
  8. Staying Engaged with Balancing Redox Equations Study Guide Pg 117
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Balancing Redox Equations Study Guide Pg 117
  9. Balancing eBooks and Physical Books Balancing Redox Equations Study Guide Pg 117
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Balancing Redox Equations Study Guide Pg 117
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Balancing Redox Equations Study Guide Pg 117
    - Setting Reading Goals Balancing Redox Equations Study Guide Pg 117
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Balancing Redox Equations Study Guide Pg 117
    - Fact-Checking eBook Content of Balancing Redox Equations Study Guide Pg 117
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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