

## Chemical Names and Formulas

### ❖ Key Vocabulary

**Monoatomic Ion** – Ions formed from a single atom

**Binary compound** – Compounds composed of two elements

**Nomenclature** – the “naming system” of binary ionic compounds

**Oxyanion** – polyatomic ions that contain oxygen

**Salt** – an ionic compound composed of a cation and the anion from an acid

### Chemical Formulas

Subscripts after an elemental symbol represent the number of atoms of the element in the compound

Molecular formulas represent how many atoms of each element are in a compound.  
( $C_2H_6$ ,  $H_2O$ ,  $C_6H_{12}O_6$ )

Ionic formulas show the simplest ratio, cations to anions, of atoms in the lattice

Parenthesis are used around a polyatomic ion in a chemical formula

Examples  $Al_2(SO_4)_3$ ,  $Cu(PO_4)_2$ ,  $Sr(NO_3)_2$

### Monoatomic Ions

– Ions formed from a single atom

Group 1 elements may lose one electron = 1+

Group 2 elements may lose two electrons = 2+

Group 13 elements may lose three electrons = 3+

Group 17 elements may gain one electron = 1-

Group 16 elements may gain two electrons = 2-

Group 15 elements may gain three electrons = 3-

Transition metals have various possibilities (no pattern)

### Naming monoatomic Ions

Monoatomic cations are simply called by the element's name

\*the stock system is used when there is more than one possibility for an element (more on this later)

Monoatomic anions end in IDE such as Fluorine – Fluoride, Oxygen – Oxide

(reference table 1 on page 221)

### Binary Ionic Compounds

1. Write the symbols for the ions with the cation first ( $Al^{3+} O^{2-}$ )

2. Cross over the charges writing the charge of one ion as the subscript for the other ion ( $Al_2^{3+} O_3^{2-}$ )

3. Multiply subscript and superscript for each ion then find the lowest whole number ratio ( $Al_2O_3$ )

**Name it!!! Aluminum Oxide**

Examples:  $Zn^{2+} I^-$

$Zn^{2+} S^{2-}$

(More on p 223 practice problems)

# Chemical Names And Formulas Study Guide

**Xiang Xie**



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## **Table of Contents Chemical Names And Formulas Study Guide**

1. Understanding the eBook Chemical Names And Formulas Study Guide
  - The Rise of Digital Reading Chemical Names And Formulas Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Names And Formulas Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Names And Formulas Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Names And Formulas Study Guide
  - Personalized Recommendations
  - Chemical Names And Formulas Study Guide User Reviews and Ratings
  - Chemical Names And Formulas Study Guide and Bestseller Lists
5. Accessing Chemical Names And Formulas Study Guide Free and Paid eBooks
  - Chemical Names And Formulas Study Guide Public Domain eBooks
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  - Chemical Names And Formulas Study Guide Budget-Friendly Options

6. Navigating Chemical Names And Formulas Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
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  - Chemical Names And Formulas Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Names And Formulas Study Guide
  - Highlighting and Note-Taking Chemical Names And Formulas Study Guide
  - Interactive Elements Chemical Names And Formulas Study Guide
8. Staying Engaged with Chemical Names And Formulas Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Names And Formulas Study Guide
9. Balancing eBooks and Physical Books Chemical Names And Formulas Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Names And Formulas Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Names And Formulas Study Guide
  - Setting Reading Goals Chemical Names And Formulas Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Names And Formulas Study Guide
  - Fact-Checking eBook Content of Chemical Names And Formulas Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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